

30102 - Chemistry

Syllabus Information

Academic Year: 2019/20

Subject: 30102 - Chemistry

Faculty / School: 175 - Escuela Universitaria Politécnica de La Almunia

179 - Centro Universitario de la Defensa - Zaragoza

Degree: 457 - Bachelor's Degree in Industrial Organisational Engineering

563 - Bachelor's Degree in Industrial Organisational Engineering

425 - Bachelor's Degree in Industrial Organisational Engineering

ECTS: 6.0

Year: 1

Semester: 425 - First semester

563 - Second semester

Subject Type: Basic Education

Module: ---

1.General information

1.1.Aims of the course

The aim of the subject is the acquisition of a basic view of the structure of matter in connection to its properties and the chemical transformations that matter can undergo.

1.2.Context and importance of this course in the degree

The achievement of the competences that are pursued with this course will allow the student to address simple practical cases related to Chemistry that may arise in their profession. In addition, the combination of the competences acquired in this subject, together with the rest of the basic training of the degree, will provide future graduates with the knowledge necessary to deal with complex cases in which different course may converge.

1.3.Recommendations to take this course

The usual ones to access the studies of any Engineering degree, basically the choice of the scientific-technological area in the Sixth Form. It is recommended to have acquired the competences of the Chemistry subject of, at least, one year in that level.

2.Learning goals

2.1.Competences

Ability to solve problems and take decisions with initiative, creativity and critical reasoning.

Ability to continue learning and develop self-learning strategies.

Ability to understand and apply the basic principles of general chemistry, organic and inorganic chemistry and its applications for engineering.

2.2.Learning goals

The student, to pass this subject, should demonstrate the following results ...

Knows the basic concepts of the structure of matter and relates them to their fundamental properties.

In a transversal way, he uses and interprets the language applied to chemical compounds and transformations.

Solve issues and problems of General Chemistry.

It uses adequate basic laboratory equipment to perform simple chemical experiments.

He interprets and presents contents from basic scientific texts related to Engineering, .

2.3.Importance of learning goals

This subject is included in the basic training module of the degree which, in a broad sense, aims to unify the knowledge of students and prepare them to address more specific subjects of the degree. In this sense, together with the rest of the basic subjects, the subject Chemistry contributes to lay the basis of a scientific model and, in addition, to equip future graduates with the necessary tools to approach other disciplines of the degree that need chemical concepts. Finally, the graduate will know and will be able to use the basic tools of Chemistry that will allow him to develop the professional competences related to this subject.

3.Assessment (1st and 2nd call)

3.1.Assessment tasks (description of tasks, marking system and assessment criteria)

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1st CALL

N is the final grade of the subject 30102 Chemistry, obtained in the evaluation of the following

- problem solving of topics 1, 2 and 3 and of practice 1
- Multiple choice questionnaire of the whole subject, including practices.

P being the note obtained in problem solving and C obtained in the questionnaire,

1. if $P \geq 4$ and $C \geq 4$, the note N is calculated as follows:

$$N = 0.30 \cdot P + 0.70 \cdot C$$

The subject is passed only if $N \geq 5$.

Students can pass the problem solving test in the partial evaluation test that will take place

2. If P and / or C are lower than 4 the student does not pass the subject and his final grade

$$\text{If } P < 4 \text{ and } C < 5 \text{ } N = 0.30 \cdot P + 0.70 \cdot C$$

$$\text{If } P < 4 \text{ and } C > 5 \text{ } N = 0.30 \cdot P + 3.5$$

$$\text{If } P < 5 \text{ and } C < 4 \text{ } N = 0.30 \cdot P + 0.70 \cdot C$$

$$\text{If } P > 5 \text{ and } C < 4 \text{ } N = 1.5 + 0.70 \cdot C$$

2nd CALL

The students that are examined in 2nd call will perform the tests described in sections a) and

Note. The scores P, C and N are calculated on 10.

4.Methodology, learning tasks, syllabus and resources

4.1.Methodological overview

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The organization of teaching will be carried out using the following steps:

- **Lectures:** Theoretical activities carried out mainly through exposition by the teacher, where the theoretical supports of the subject are displayed, highlighting the fundamental, structuring them in topics and or sections, interrelating them.
- **Practice Sessions:** The teacher resolves practical problems or cases for demonstrative purposes. This type of teaching complements the theory shown in the lectures with practical aspects.
- **Laboratory Workshop:** The lecture group is divided up into various groups, according to the number of registered students, but never with more than 16 students, in order to make up smaller sized groups.
- **Individual Tutorials:** Those carried out giving individual, personalized attention with a teacher from the department. Said tutorials may be in person or online.

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In order to get the objectives described in the previous sections the following learning activities will be carried out: lectures, practical and problem solving sessions, laboratory sessions.

4.2. Learning tasks

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The subject has 6 ECTS credits, which represents 150 hours of student work in the subject during the trimester, in other words, 10 hours per week for 15 weeks of class.

A summary of a weekly timetable guide can be seen in the following table. These figures are obtained from the subject file in the Accreditation Report of the degree, taking into account the level of experimentation considered for the said subject is moderate.

Activity	Weekly school hours
Lectures	2
Laboratory Workshop	2
Other Activities	6

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1. lectures,
2. practical and problem-solving sessions,
3. laboratory sessions.

4.3. Syllabus

The course will address the following topics:

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THEORETICAL CONTENTS

Module	Contents
1. THE ATOM AND THE PERIODIC SYSTEM	<p>Topic 1.- The atom. Elemental particles. Atomic models. The Bohr atom. The quantum mechanical model. Atomic orbitals; quantum numbers. Principles for the electronic construction of atoms.</p> <p>Topic 2.- General overview of the periodic table Description of the current periodic table: groups and periods. Study of the electron shell and the periodic system. Periodic properties.</p>
2. THE CHEMICAL BOND	<p>Topic 3.- The ionic bond General characteristics of the ionic bond. Network energy. General properties of ionic compounds.</p> <p>Topic 4.- The covalent bond Simplified model: the Lewis theory. Bond polarity and geometry. Valence bond theory. Orbital hybridization. Molecular orbital theory.</p> <p>Topic 5.- The metallic bond General characteristics of metals. Theories of the metallic bond: the electron sea theory and valence bond theory. Alloys: classes.</p>
3. BONDS BETWEEN MOLECULES	<p>Topic 6.- Intermolecular bonds Van der Waal forces. Hydrogen bonds.</p>

4. AGGREGATION STATES	<p>Topic 7.- The gas state General characteristics of gases. Laws that govern the gas state. Equations of state. Kinetic theory. Gas mixtures: Dalton's Law. Gas diffusion and effusion: Graham's Law. Real gases: The Van der Waal equation.</p> <p>Topic 8.- The liquid state General characteristics of liquids. Vapour pressure. The effect of temperature on vapour pressure. Critical phenomena. Condensation of vapours and gases. Solidification.</p> <p>Topic 9.- The solid state Characteristics of solids. Classes of crystal network. Classes of solids based on bonding type. The phase rule and the triple point.</p>
5. INTRODUCTION TO THE STUDY OF SOLUTIONS	<p>Topic 10.- Introduction to the study of solutions Disperse systems. Types of solutions. Means of expressing concentration. Solid-in-liquid solutions. Liquid-in-liquid solutions. Gas-in-liquid solutions. Colligative properties of solutions. Colloidal solutions.</p>
6. INTRODUCTION TO THE STUDY OF REACTIONS	<p>Topic 11.- Chemical equilibrium The concept of reaction rate. Reversible and irreversible reactions. Chemical equilibrium: the equilibrium constant. Le Chatelier's principle. Stable, unstable and metastable systems.</p> <p>Topic 12.- Neutralisation reactions The acid-base concept. Aqueous solutions: pH of aqueous solutions. Acid-base strengths. Equilibrium constants. Salt hydrolysis.</p>
7. ORGANIC CHEMISTRY	<p>Topic 13.- Organic chemistry The properties of carbon. Types of organic substances Isomers. Organic chemical reactions. Polymers</p>

PRACTICAL CONTENTS

Practical 1	Standards in Chemical Laboratory <i>Techniques, Equipment and Safety</i>
Practical 2	Solution Preparation <i>Na₂CO₃ 0,1 M from Na₂CO₃ solid; CaCl₂ 0,1 M from CaCl₂ 2 M</i>
Practical 3	Filtration <i>Gravity Filtration and Vacuum Filtration</i>
Practical 4	Volumetric Analysis <i>Water hardness; carbonates and bicarbonates in water.</i>
Practical 5	Distillation

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Contents

Unit 1. Chemistry and matter

- 1.1. Basic concepts
- 1.2. Fundamental laws of chemical reactions. Atoms
- 1.3. Compounds and chemical formula
- 1.4. Composition of compounds
- 1.5. Mixtures, solutions and solubility

Unit 2. Chemical reactions and stoichiometry

- 2.1. Chemical reaction and chemical equation
 - 2.2. Stoichiometry of chemical reactions
 - 2.3. Gases
 - 2.3. Aqueous reactions
- Lab. session 1: Preparation of Lead(II) Iodide.

Unit 3. Chemical equilibrium

- 3.1. Basic concepts
- 3.2. Equilibrium constant and chemical equation
- 3.3. Numerical importance of the equilibrium constant
- 3.4. The reaction quotient, Q
- 3.5. Le Chatelier's principle
- 3.6. Acid-base equilibria and the pH scale

Unit 4. Energy and Chemistry

- 4.1. Basic concepts
- 4.2. The First Law of Thermodynamics: heat, work and enthalpy
- 4.3. Spontaneous processes
- 4.4. Entropy.
- 4.6. The Second Law of Thermodynamics.

Unit 5. Chemical kinetics

- 5.1. Reaction rates
 - 5.2. Concentration and reaction rate
 - 5.3. Rate laws
 - 5.4. The change of concentration with time.
 - 5.5. Temperature and reaction rate
 - 5.6. Reaction mechanisms
 - 5.7. Catalysis
- Lab. session 2: The aluminothermic reaction

Unit 6. Atomic models

- 6.1. From classical physics to quantum mechanics
- 6.2. Line spectra and the Bohr model
- 6.3. The wave behaviour of matter
- 6.4. Quantum mechanics and atomic orbitals
- 6.5. Many-electron atoms
- 6.6. Electronic configurations and the Periodic Table
- 6.7. Periodic properties

Unit 7. Chemical bond I

- 7.1. Chemical bonds, Lewis symbols and the octet rule
- 7.2. Ionic bonding
- 7.3. Covalent bonding
- 7.4. Strength of covalent bonds

Unit 8. Chemical bond II

- 8.1. The VSEPR model
- 8.2. Molecular polarity
- 8.3. Molecular-orbital model
- 8.4. Molecular-orbital model for metals

Unit 9. Intermolecular forces and liquids

- 9.1. Intermolecular forces
- 9.2. Solubility and the solution process
- 9.3. Phases, phase changes and phase diagrams
- 9.4. Liquids: boiling point, vapor pressure, surface tension, viscosity

Unit 10. Solids

- 10.1. Amorphous solids and the short-range order. Glass transition temperature
- 10.2. Crystalline solids, the long-range order and the unit cell
- 10.3. Bonding in solids

Unit 11. Materials

- 11.1. Materials science and classes of materials

- 11.2. Ceramics
- 11.3. Polymers
- 11.4. Alloys
- 11.5. Semiconductors

4.4. Course planning and calendar

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SECTION	Topic	Nº hours
0 y 1	Presentation. Atom (<i>Topic 1</i>) and the Periodic System (<i>Topic 2</i>)	10
2 y 3	The Chemical Bond (<i>Topics 3, 4 y 5</i>) and Intermolecular Bonds (<i>Topic 6</i>)	12
4	Aggregation States (<i>Topics 7, 8 y 9</i>)	6
5	Introduction to the study of solutions (<i>Topic 10</i>)	8
6	Introduction to the study of reactions (<i>Topics 11 y 12</i>)	8
7	Organic Chemistry (<i>Topic 13</i>)	4
	Practical Course	6
	Exams	6
TOTAL		60

The dates of the final exams will be those that are officially published at <http://www.eupla.unizar.es/asuntos-academicos/examenes>

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Before the start of the semester the course additional information will be published in the Moodle platform, which can be consulted at <http://moodle.unizar.es> after authentication with the student's username and password.

This information will include the course planning, materials, bibliography and other recommendations to follow the course.

Information about general course calendars and timetables can also be found at the website of the Centro Universitario de la Defensa: <http://cud.unizar.es>.

4.5. Bibliography and recommended resources

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REFERENCES AND RESOURCES

Basic material prepared by the lecturer:

GONZÁLEZ PAÚLES, J. y BURBANO GARCÍA, G. *Apuntes de Química para el Grado en Ingeniería Civil*. Ed. Eupla. 2011. ISBN 978 84 694 0618 2

GONZÁLEZ PAÚLES, J. *Química General. Apuntes del curso práctico*. Ed. Eupla. 2010. ISBN 978-84-692-8044-7

Power Point in Platform Moodle (<https://moodle2.unizar.es/>)

BIBLIOGRAPHY

The bibliography can be consulted on webpage

http://biblos.unizar.es/br/br_citas.php?codigo=30102&year=2019

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RESOURCES AND MATERIAL

Power Point Slides: The subject is presented with the help of Power Point software. The pdf files related to each of the topics will be available to students on the Moodle platform (<https://moodle.unizar.es/>)

Chemical Nomenclature:

- Principles

- Test

Material related to the Laboratory practices:

- Introduction to the work in lab.

- Theoretical and experimental abstracts of practices.

Audio-visual Material:

. Videos to illustrate different aspects of the subject syllabus.

Bibliography:

http://biblos.unizar.es/br/br_citas.php?codigo=30102&year=2019

Any other additional material used will be deposited in the Moodle platform.